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(Acetylacetone isonicotinoylhydrazonato-κ^3O,N',O')dioxidovanadate(V) monohydrate

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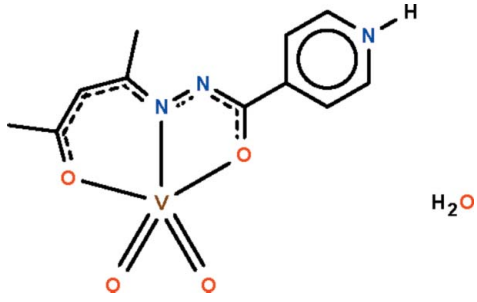
Key indicators: single-crystal X-ray study; T = 100 K; mean σ(C–C) = 0.004 Å; R factor = 0.039; wR factor = 0.121; data-to-parameter ratio = 12.9.

The hydrazone anion in the title compound, $[V(C_{11}H_{12}N_3O_2)O_2]\cdot H_2O$, is zwitterionic as its pyridyl N atom is protonated; the O, N and O′ atoms span the axial–equatorial–axial positions of the trigonal-bipyramidal coordination polyhedron of the metal atom. All non-H atoms lie on a crystallographic mirror plane apart from the oxide ligands, which are related by mirror symmetry. The pyridinium N atom acts as a hydrogen-bond donor to the solvent water molecule, which is in turn a hydrogen-bond donor to the both oxide ligands. These hydrogen-bonding interactions give rise to a three-dimensional network motif.

Related literature

For related vanadium(V) structures, see: Shao *et al.* (1988). The reaction of oxidovanadium(IV) bis(acetylacetonate), $VO(acac)_2$, with aroylhydrazines in methanol yields Schiff-base complexes having the dinuclear $[V(=O)(\mu\text{-OMe})_2V(=O)]^{4+}$ core, see: Sarkari & Pal (2009).



Experimental

Crystal data

$[V(C_{11}H_{12}N_3O_2)O_2]\cdot H_2O$
M_r = 319.19
Orthorhombic, *Pnma*
a = 13.9848 (10) Å
b = 6.6630 (4) Å
c = 13.8904 (10) Å
V = 1294.32 (15) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 0.79 mm^{-1}
T = 100 K
0.35 × 0.20 × 0.20 mm

Data collection

Bruker SMART APEX diffractometer
Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)
T_{min} = 0.770, T_{max} = 0.858
11995 measured reflections
1610 independent reflections
1416 reflections with $I > 2\sigma(I)$
R_{int} = 0.033

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.039
$wR(F^2)$ = 0.121
S = 1.11
1610 reflections
125 parameters
2 restraints
H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{max}$ = 0.75 e Å^{-3}
$\Delta\rho_{min}$ = −0.72 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
O1W—H1w···O1	0.84 (1)	1.91 (1)	2.732 (2)	168 (3)
N3—H3···O1w^{i}	0.86 (1)	1.87 (3)	2.683 (4)	158 (6)

Symmetry code: (i) $-x+\frac{3}{2}, -y, z+\frac{1}{2}$.

Data collection: *APEX2* (Bruker, 2009); cell refinement: *SAINT* (Bruker, 2009); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *X-SEED* (Barbour, 2001); software used to prepare material for publication: *publCIF* (Westrip, 2010).

We thank the University of Malaya (RG020/09AFR) for supporting this study.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: NK2049).

References

Barbour, L. J. (2001). *J. Supramol. Chem.* **1**, 189–191.
Bruker (2009). *APEX2* and *SAINT*. Bruker AXS Inc., Madison, Wisconsin, USA.
Sarkari, A. & Pal, S. (2009). *Inorg. Chim. Acta*, **362**, 3807–3812.
Shao, M.-C., Zhang, Y.-J., Zhang, Z.-Y. & Tang, Y.-Q. (1988). *Sci. Chin. Ser. B (Engl. Ed.)*, **31**, 781–788.
Sheldrick, G. M. (1996). *SADABS*. University of Göttingen, Germany.
Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.
Westrip, S. P. (2010). *J. Appl. Cryst.* **43**, 920–925.

supplementary materials

Acta Cryst. (2010). E**66**, m1020 [**doi:10.1107/S1600536810028886**]

(Acetylacetone isonicotinoylhydrazonato-κ^3O,N',O')dioxidovanadate(V) monohydrate

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Comment

The reaction of oxovanadium(IV) bis(acetylacetonate), $VO(acac)_2$, with aroylhydrazines in acetonitrile yields vanadium(V) compounds of the formulation $V_2O_3L_2$ (where *L* represents the doubly-deprotonated Schiff base). In methanol, the reaction yields Schiff-base complexes having the dinuclear $[V(=O)(\mu\text{-}OMe)_2V(=O)]^{4+}$ core (Sarkari & Pal, 2009). In the present study, the reaction with isonicotinic acid hydrazide yields the expected vanadium(V) complex of the mono-deprotonated Schiff base as a negatively-charged zwitterion as the pyridyl N-atom is protonated (Scheme I). The metal atom shows trigonal bipyramidal coordination, with the *O*,*N*,*O*'-atoms of the Schiff base spanning the axial sites (Fig. 1).

All non-hydrogen atoms lie on a crystallographic mirror plane other than the oxo ligands, which are related by mirror symmetry. The pyridinium N atom acts as a hydrogen-bond donor to the solvate water molecule, which is in turn a hydrogen bond donor to the both oxo ligands. Hydrogen bonding gives rise to a three-dimensional network motif.

Experimental

Bis(acetylacetonato)oxovanadium(IV) (0.13 g, 0.5 mmol) and isonicotinic acid hydrazide (0.07 g, 0.75 mmol) heated in methanol (50 ml) for one hour. The brown solution was filtered; slow evaporation of the filtrate afforded brown crystals.

Refinement

Carbon-bound H-atoms were placed in calculated positions (C—H 0.95 to 0.98 Å) and were included in the refinement in the riding model approximation, with *U*(H) set to 1.2 to 1.5*U*(C). The methyl carbons lies on a mirror plane, so that one of the H atoms lies on the plane whereas the other lies on a general position.

The amino and water H-atoms were located in a difference Fourier map, and were refined with distance restraints of N–H 0.86±0.01 and O–H 0.84±0.01 Å; their temperature factors were freely refined.

Figures

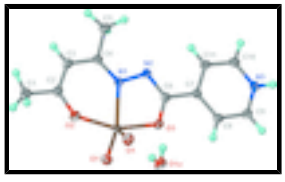

Fig. 1. Thermal ellipsoid plot (Barbour, 2001) of $VO_2(C_{11}H_{12}N_3O_2)\cdot H_2O$ at the 70% probability level; hydrogen atoms are drawn as spheres of arbitrary radius. Symmetry transformation: i = *x*, 1/2 – *y*, *z*.

(Acetylacetone isonicotinoylhydrazonato- κ^3O,N',O')dioxidovanadate(V) monohydrate

Crystal data

$[V(C_{11}H_{12}N_3O_2)O_2]\cdot H_2O$	$F(000) = 656$
$M_r = 319.19$	$D_x = 1.638$ Mg m^{-3}
Orthorhombic, *Pnma*	Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
Hall symbol: -P 2ac 2n	Cell parameters from 3731 reflections
$a = 13.9848$ (10) Å	$\theta = 2.9$–$27.6°$
$b = 6.6630$ (4) Å	$\mu = 0.79$ mm^{-1}
$c = 13.8904$ (10) Å	$T = 100$ K
$V = 1294.32$ (15) Å^3	Prism, brown
$Z = 4$	0.35 × 0.20 × 0.20 mm

Data collection

Bruker SMART APEX diffractometer	1610 independent reflections
Radiation source: fine-focus sealed tube	1416 reflections with $I > 2\sigma(I)$
graphite	$R_{int} = 0.033$
ω scans	$\theta_{max} = 27.5°$, $\theta_{min} = 2.1°$
Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)	$h = -18\rightarrow18$
$T_{min} = 0.770$, $T_{max} = 0.858$	$k = -8\rightarrow8$
11995 measured reflections	$l = -17\rightarrow18$

Refinement

Refinement on F^2	Primary atom site location: structure-invariant direct methods
Least-squares matrix: full	Secondary atom site location: difference Fourier map
$R[F^2 > 2\sigma(F^2)] = 0.039$	Hydrogen site location: inferred from neighbouring sites
$wR(F^2) = 0.121$	H atoms treated by a mixture of independent and constrained refinement
$S = 1.11$	$w = 1/[\sigma^2(F_o^2) + (0.0616P)^2 + 2.2562P]$ where $P = (F_o^2 + 2F_c^2)/3$
1610 reflections	$(\Delta/\sigma)_{max} = 0.001$
125 parameters	$\Delta\rho_{max} = 0.75$ e Å^{-3}
2 restraints	$\Delta\rho_{min} = -0.72$ e Å^{-3}

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}	Occ. (<1)
V1	0.41194 (4)	0.2500	0.31526 (4)	0.01205 (19)	
O1	0.41890 (11)	0.0504 (3)	0.24747 (12)	0.0191 (4)	

O2	0.27403 (16)	0.2500	0.32874 (16)	0.0204 (5)	
O3	0.54634 (16)	0.2500	0.35590 (16)	0.0189 (5)	
O1W	0.47876 (17)	−0.2500	0.12798 (18)	0.0194 (5)	
H1W	0.467 (2)	−0.148 (3)	0.161 (2)	0.040 (10)*	
N1	0.40746 (18)	0.2500	0.46744 (19)	0.0138 (5)	
N2	0.49679 (19)	0.2500	0.51358 (19)	0.0139 (5)	
N3	0.84943 (19)	0.2500	0.5420 (2)	0.0155 (5)	
H3	0.9099 (11)	0.2500	0.554 (4)	0.059 (18)*	
C1	0.1102 (2)	0.2500	0.3685 (3)	0.0218 (7)	
H1A	0.1065	0.2500	0.2981	0.033*	
H1B	0.0784	0.3701	0.3937	0.033*	0.50
H1C	0.0784	0.1299	0.3937	0.033*	0.50
C2	0.2135 (2)	0.2500	0.3993 (2)	0.0163 (6)	
C3	0.2387 (2)	0.2500	0.4946 (2)	0.0165 (6)	
H3A	0.1889	0.2500	0.5412	0.020*	
C4	0.3340 (2)	0.2500	0.5277 (2)	0.0137 (6)	
C5	0.3508 (2)	0.2500	0.6351 (2)	0.0198 (7)	
H5A	0.4197	0.2500	0.6480	0.030*	
H5B	0.3218	0.1299	0.6635	0.030*	0.50
H5C	0.3218	0.3701	0.6635	0.030*	0.50
C6	0.5624 (2)	0.2500	0.4488 (2)	0.0140 (6)	
C7	0.6634 (2)	0.2500	0.4810 (2)	0.0133 (6)	
C8	0.7378 (2)	0.2500	0.4146 (2)	0.0156 (6)	
H8	0.7246	0.2500	0.3475	0.019*	
C9	0.8314 (2)	0.2500	0.4473 (2)	0.0165 (6)	
H9	0.8828	0.2500	0.4025	0.020*	
C10	0.7787 (2)	0.2500	0.6080 (2)	0.0166 (6)	
H10	0.7941	0.2500	0.6746	0.020*	
C11	0.6845 (2)	0.2500	0.5800 (2)	0.0153 (6)	
H11	0.6347	0.2500	0.6265	0.018*	

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
V1	0.0111 (3)	0.0163 (3)	0.0087 (3)	0.000	0.00041 (18)	0.000
O1	0.0180 (8)	0.0203 (8)	0.0192 (8)	0.0007 (6)	−0.0002 (6)	−0.0048 (7)
O2	0.0131 (11)	0.0351 (14)	0.0129 (11)	0.000	−0.0001 (9)	0.000
O3	0.0127 (11)	0.0333 (14)	0.0108 (11)	0.000	−0.0004 (8)	0.000
O1W	0.0188 (12)	0.0193 (12)	0.0201 (12)	0.000	0.0066 (10)	0.000
N1	0.0109 (12)	0.0183 (13)	0.0121 (13)	0.000	−0.0012 (9)	0.000
N2	0.0121 (12)	0.0171 (12)	0.0124 (12)	0.000	−0.0019 (10)	0.000
N3	0.0126 (12)	0.0192 (13)	0.0146 (13)	0.000	−0.0021 (10)	0.000
C1	0.0122 (15)	0.035 (2)	0.0187 (16)	0.000	−0.0006 (13)	0.000
C2	0.0128 (14)	0.0195 (15)	0.0165 (16)	0.000	0.0011 (12)	0.000
C3	0.0144 (15)	0.0194 (15)	0.0158 (15)	0.000	0.0010 (12)	0.000
C4	0.0154 (15)	0.0141 (14)	0.0115 (14)	0.000	0.0005 (11)	0.000
C5	0.0178 (15)	0.0314 (19)	0.0102 (15)	0.000	0.0025 (12)	0.000
C6	0.0143 (14)	0.0150 (14)	0.0126 (14)	0.000	−0.0013 (12)	0.000

C7	0.0140 (14)	0.0131 (14)	0.0128 (14)	0.000	−0.0016 (11)	0.000
C8	0.0170 (15)	0.0180 (15)	0.0117 (14)	0.000	−0.0009 (12)	0.000
C9	0.0154 (15)	0.0195 (15)	0.0145 (15)	0.000	0.0010 (12)	0.000
C10	0.0185 (15)	0.0191 (15)	0.0123 (15)	0.000	−0.0015 (12)	0.000
C11	0.0156 (15)	0.0181 (15)	0.0122 (14)	0.000	0.0021 (12)	0.000

Geometric parameters (Å, °)

V1—O1	1.6323 (17)	C1—H1C	0.9800
V1—O1^{i}	1.6323 (17)	C2—C3	1.369 (5)
V1—O2	1.938 (2)	C3—C4	1.411 (4)
V1—O3	1.962 (2)	C3—H3A	0.9500
V1—N1	2.115 (3)	C4—C5	1.510 (4)
O2—C2	1.295 (4)	C5—H5A	0.9800
O3—C6	1.310 (4)	C5—H5B	0.9800
O1W—H1W	0.838 (10)	C5—H5C	0.9800
N1—C4	1.325 (4)	C6—C7	1.481 (4)
N1—N2	1.404 (4)	C7—C8	1.391 (4)
N2—C6	1.286 (4)	C7—C11	1.407 (4)
N3—C9	1.340 (4)	C8—C9	1.385 (4)
N3—C10	1.349 (4)	C8—H8	0.9500
N3—H3	0.861 (10)	C9—H9	0.9500
C1—C2	1.507 (4)	C10—C11	1.373 (5)
C1—H1A	0.9800	C10—H10	0.9500
C1—H1B	0.9800	C11—H11	0.9500
O1—V1—O1^{i}	109.11 (13)	C2—C3—H3A	118.1
O1—V1—O2	96.61 (7)	C4—C3—H3A	118.1
O1^{i}—V1—O2	96.61 (7)	N1—C4—C3	121.8 (3)
O1—V1—O3	96.25 (7)	N1—C4—C5	120.3 (3)
O1^{i}—V1—O3	96.25 (7)	C3—C4—C5	117.9 (3)
O2—V1—O3	157.73 (10)	C4—C5—H5A	109.5
O1—V1—N1	125.34 (6)	C4—C5—H5B	109.5
O1^{i}—V1—N1	125.34 (6)	H5A—C5—H5B	109.5
O2—V1—N1	82.75 (10)	C4—C5—H5C	109.5
O3—V1—N1	74.98 (10)	H5A—C5—H5C	109.5
C2—O2—V1	136.3 (2)	H5B—C5—H5C	109.5
C6—O3—V1	116.6 (2)	N2—C6—O3	124.5 (3)
C4—N1—N2	113.6 (3)	N2—C6—C7	118.0 (3)
C4—N1—V1	130.9 (2)	O3—C6—C7	117.5 (3)
N2—N1—V1	115.46 (19)	C8—C7—C11	119.4 (3)
C6—N2—N1	108.4 (3)	C8—C7—C6	120.9 (3)
C9—N3—C10	122.0 (3)	C11—C7—C6	119.7 (3)
C9—N3—H3	112 (4)	C9—C8—C7	119.3 (3)
C10—N3—H3	126 (4)	C9—C8—H8	120.3
C2—C1—H1A	109.5	C7—C8—H8	120.3
C2—C1—H1B	109.5	N3—C9—C8	120.0 (3)
H1A—C1—H1B	109.5	N3—C9—H9	120.0
C2—C1—H1C	109.5	C8—C9—H9	120.0

H1A—C1—H1C	109.5	N3—C10—C11	120.7 (3)
H1B—C1—H1C	109.5	N3—C10—H10	119.7
O2—C2—C3	124.3 (3)	C11—C10—H10	119.7
O2—C2—C1	114.3 (3)	C10—C11—C7	118.6 (3)
C3—C2—C1	121.4 (3)	C10—C11—H11	120.7
C2—C3—C4	123.9 (3)	C7—C11—H11	120.7
O1—V1—O2—C2	124.90 (6)	N2—N1—C4—C3	180.0
O1^{i}—V1—O2—C2	−124.90 (6)	V1—N1—C4—C3	0.0
O3—V1—O2—C2	0.0	N2—N1—C4—C5	0.0
N1—V1—O2—C2	0.0	V1—N1—C4—C5	180.0
O1—V1—O3—C6	−124.96 (6)	C2—C3—C4—N1	0.0
O1^{i}—V1—O3—C6	124.96 (6)	C2—C3—C4—C5	180.0
O2—V1—O3—C6	0.0	N1—N2—C6—O3	0.0
N1—V1—O3—C6	0.0	N1—N2—C6—C7	180.0
O1—V1—N1—C4	−92.98 (9)	V1—O3—C6—N2	0.0
O1^{i}—V1—N1—C4	92.98 (9)	V1—O3—C6—C7	180.0
O2—V1—N1—C4	0.0	N2—C6—C7—C8	180.0
O3—V1—N1—C4	180.0	O3—C6—C7—C8	0.0
O1—V1—N1—N2	87.02 (9)	N2—C6—C7—C11	0.0
O1^{i}—V1—N1—N2	−87.02 (9)	O3—C6—C7—C11	180.0
O2—V1—N1—N2	180.0	C11—C7—C8—C9	0.0
O3—V1—N1—N2	0.0	C6—C7—C8—C9	180.0
C4—N1—N2—C6	180.0	C10—N3—C9—C8	0.0
V1—N1—N2—C6	0.0	C7—C8—C9—N3	0.0
V1—O2—C2—C3	0.0	C9—N3—C10—C11	0.0
V1—O2—C2—C1	180.0	N3—C10—C11—C7	0.0
O2—C2—C3—C4	0.0	C8—C7—C11—C10	0.0
C1—C2—C3—C4	180.0	C6—C7—C11—C10	180.0

Symmetry codes: (i) *x*, −*y*+1/2, *z*.

Hydrogen-bond geometry (Å, °)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
O1W—H1w···O1	0.84 (1)	1.91 (1)	2.732 (2)	168 (3)
N3—H3···O1w^{ii}	0.86 (1)	1.87 (3)	2.683 (4)	158 (6)

Symmetry codes: (ii) −*x*+3/2, −*y*, *z*+1/2.

Fig. 1